

## Regents Questions On Electrochemistry

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Unit 12: Electrochemistry-Key Regents Chemistry '14-'15 Mr. Murdoch Page 13 of 61 Key e. Solving for X, we get +6, so the oxidation number for sulfur in the polyatomic ion (SO 4-2) is +6, and so sulfur will be +6 in ZnSO 4. \*Check the Periodic Table to see if sulfur has a +6 charge listed; it does!

### Unit 12: Electrochemistry-Key Regents Chemistry '14 Mr ...

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### Regents Questions On Electrochemistry

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Oxidation: loss of electrons Reduction: gain of electrons Oxidizing Agent: substance reduced Reducing Agent: substance oxidized Anode: electrode where oxidation occurs (AN OX) Cathode: electrode where reduction occurs (RED CAT) Salt Bridge: allows the flow of ions between the anode and cathode solutions to keep a balance of charge Cations flow toward the Cathode

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Electrochemistry is the chemistry of electrochemical reactions; which deal with the relationship between electrical energy and chemical reactions. Electrochemical reactions involve transfer of electrons and are essentially REDOX reactions. Displacement and REDOX Reactions Experiment 1:- Displacement Reactions among Metals. Procedure

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PRACTICE PACKET: ELECTROCHEMISTRY 3 • All Free elements are zero • Group 1 is +1 • Group 2 is +2 • Fluorine is -1 • H is usually +1, and if so, will be first written in the chemical formula. • O is almost always -2, exceptions: H<sub>2</sub>O 2 and OF<sub>2</sub> • In compounds using polyatomic ions, the first and middle elements all have + oxidation states. • The sum of oxidation states in a ...

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