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combined gas law, is a combination of all the variables in the previous gas laws. The ideal gas law is expressed by the formula $PV = nRT$ where $P =$ pressure $V =$ volume $n =$ number of moles of gas $R =$ ideal gas constant $T =$ absolute temperature The value of R depends on the units of pressure, volume and temperature.

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STUDY GUIDE AP Chemistry Chapter
Ten- Gases Lecture Notes 10.1

Characteristics of Gases •All substances
have three phases: solid, liquid and gas.
•Substances that are liquids or solids
under ordinary conditions may also exist

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as gases. •These are often referred to as vapors. •Many of the properties of gases differ from those of solids and liquids:

Chapter Ten- Gases #2 Pg 432 #5, 43, 45, 47, #3 Pg 432 #6 ...

The pressure of a gas is directly proportional to the temperature if the volume remains constant What is the

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formula for the combined gas law?

$$P_1V_1/T_1=P_2V_2/T_2$$

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190 Study Guide for An Introduction to
Chemistry Section Goals and
Introductions Section 13.1 Gases and
Their Properties Goals To describe the

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particle nature of both real and ideal gases. To describe the properties of gases that can be used to explain their characteristics: volume, number of particles, temperature, and pressure.

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$P_1V_1=P_2V_2$. - Charlie's law states that the
volume of a fixed amount of gas is

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directly proportional to its kelvin temperature at constant pressure.

$V_1/T_1 = V_2/T_2$. -Gay - Lussac's law states that the pressure of a fixed amount of gas is directly proportional to its kelvin temperature at constant value.

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Ideal Gas Law and Stoichiometry Use the following reaction to answer the next few questions: $2 \text{C}_8\text{H}_{18}(\text{l}) + 25 \text{O}_2(\text{g}) \dots > 16 \text{CO}_2(\text{g}) + 18 \text{H}_2\text{O}(\text{g})$ The above reaction is the reaction between gasoline (octane) and oxygen that occurs inside automobile engines.

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The solubility of gases increases with increases pressure. When a can of soda is opened, the pressure is lowered, decreasing the solubility of carbon dioxide. This causes bubbles of carbon dioxide to come out of solution. 8.

Solids, Liquids, and Gases

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Questions | Shmoop

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