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range of the nuclear  
force allows them to  
attract only protons  
very close to them, as  
shown in Figure 22-3.  
So as the number of

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protons in a nucleus increases, the electrostatic force between protons increases faster than the nuclear force. More neutrons are required to increase the nuclear force and stabilize the nucleus.

## **CHAPTER 22 Nuclear Chemistry**

Table 22.1 and example 22.1, problems 22.1-2 Beta emission (occurs when

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an atom has too many neutrons, or not enough protons). In the nucleus a neutron splits into an electron and a proton ( $1\ 0\ n \rightarrow 1\ 1\ p + 0\ -1\ e$ ) Since this electron originated from the nucleus, it is a beta particle.

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set (45) Radioactive  
decay.

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fusion, controlled nuclear reactions, radiation. Section 22.1

Define nucleon, isotope, nuclide, and nuclear reaction.

Summarize the differences between nuclear reactions and chemical reactions.

Section 22.2 Define radioactivity and radionuclide. Write balanced equations for nuclear reactions,

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Chapter 22: Nuclear  
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22-1: The Nucleus •

Atomic nuclei= protons  
and neutrons (together  
are nucleons) o

Nuclide= an  
atom—identified by #  
of protons/neutrons in  
nucleus Mass Defect  
and Nuclear Stability •

Mass defect=  
difference between

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mass of an atom and  
sum of the masses of p  
rotons/neutrons/electro  
ns o Caused by  
conversion of mass to  
energy when nucleus  
forms Nuclear ...

### **Chapter 22 Notes - Chapter 22 Nuclear Chemistry Section 22 ...**

1 CK-12 Physics

Concepts -

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Key Chapter 22: The

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Review Questions 1.

What do different isotopes of a given element have in common? How are they different? 2. What is the element name of the atom represented by? 3. How many protons and how many neutrons does the nucleus in problem #2 have?

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The \$1.3 billion bailout  
of two Ohio nuclear  
plants was already law.  
Yet, in the weeks

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before then-House

Speaker Larry

Householder's arrest,

the energy company at

the center of it all

donated ...

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